

PCTWORLD INTELLECTUAL PROPERTY ORGANIZATION
International Bureau**BU**

INTERNATIONAL APPLICATION PUBLISHED UNDER THE PATENT COOPERATION TREATY (PCT)

(51) International Patent Classification ⁶ : A01N 59/00, 25/22, C02F 1/76	A1	(11) International Publication Number: WO 99/62339 (43) International Publication Date: 9 December 1999 (09.12.99)
(21) International Application Number: PCT/US99/09801 (22) International Filing Date: 3 May 1999 (03.05.99) (30) Priority Data: 09/088,300 1 June 1998 (01.06.98) US (71) Applicant: ALBEMARLE CORPORATION [US/US]; 451 Florida Street, Baton Rouge, LA 70801-1765 (US). (72) Inventors: MOORE, Robert, M., Jr.; 6152 Belle Grove Drive, Baton Rouge, LA 70820 (US). NALEPA, Christopher, J.; 4142 Country Hill Drive, Baton Rouge, LA 70816 (US). (74) Agents: PIPPENGER, Philip, M. et al.; Albemarle Corporation, Patent & Trademark Division, 451 Florida Street, Baton Rouge, LA 70801-1765 (US).	(81) Designated States: AU, CA, JP, European patent (AT, BE, CH, CY, DE, DK, ES, FI, FR, GB, GR, IE, IT, LU, MC, NL, PT, SE). Published <i>With international search report.</i>	
(54) Title: CONCENTRATED AQUEOUS BROMINE SOLUTIONS AND THEIR PREPARATION (57) Abstract Described is a process of producing a concentrated liquid biocide formulation. Mixed together are (a) bromine chloride or bromine and (b) an aqueous solution of alkali metal salt of sulfamic acid having a pH of at least about 7, in amounts such that (i) the active bromine content of the solution is at least about 100,000 ppm (wt/wt), and (ii) the atom ratio of nitrogen to active bromine from (a) and (b) is greater than 1 when bromine is used and is greater than 0.93 when bromine chloride is used. Use of bromine chloride as the source of the active bromine in the process is preferred because in the resulting aqueous compositions, all of the bromine chloride is made available as active bromine in solution. In other words, the chlorine of the bromine chloride is converted in the process to dissolved alkali metal chloride salt, thereby liberating all of the bromine as the active bromine content of the biocidal composition.		

FOR THE PURPOSES OF INFORMATION ONLY

Codes used to identify States party to the PCT on the front pages of pamphlets publishing international applications under the PCT.

AL	Albania	ES	Spain	LS	Lesotho	SI	Slovenia
AM	Armenia	FI	Finland	LT	Lithuania	SK	Slovakia
AT	Austria	FR	France	LU	Luxembourg	SN	Senegal
AU	Australia	GA	Gabon	LV	Latvia	SZ	Swaziland
AZ	Azerbaijan	GB	United Kingdom	MC	Monaco	TD	Chad
BA	Bosnia and Herzegovina	GE	Georgia	MD	Republic of Moldova	TG	Togo
BB	Barbados	GH	Ghana	MG	Madagascar	TJ	Tajikistan
BE	Belgium	GN	Guinea	MK	The former Yugoslav Republic of Macedonia	TM	Turkmenistan
BF	Burkina Faso	GR	Greece	ML	Mali	TR	Turkey
BG	Bulgaria	HU	Hungary	MN	Mongolia	TT	Trinidad and Tobago
BJ	Benin	IE	Ireland	MR	Mauritania	UA	Ukraine
BR	Brazil	IL	Israel	MW	Malawi	UG	Uganda
BY	Belarus	IS	Iceland	MX	Mexico	US	United States of America
CA	Canada	IT	Italy	NE	Niger	UZ	Uzbekistan
CF	Central African Republic	JP	Japan	NL	Netherlands	VN	Viet Nam
CG	Congo	KE	Kenya	NO	Norway	YU	Yugoslavia
CH	Switzerland	KG	Kyrgyzstan	NZ	New Zealand	ZW	Zimbabwe
CI	Côte d'Ivoire	KP	Democratic People's Republic of Korea	PL	Poland		
CM	Cameroon	KR	Republic of Korea	PT	Portugal		
CN	China	KZ	Kazakhstan	RO	Romania		
CU	Cuba	LC	Saint Lucia	RU	Russian Federation		
CZ	Czech Republic	LI	Liechtenstein	SD	Sudan		
DE	Germany	LK	Sri Lanka	SE	Sweden		
DK	Denmark	LR	Liberia	SG	Singapore		
EE	Estonia						

CONCENTRATED AQUEOUS BROMINE SOLUTIONS AND THEIR PREPARATION

BACKGROUND

Bromine-based biocides have proven biocidal advantages over chlorination-dechlorination for the microbiological control of cooling waters and disinfection of waste treatment systems. The water treatment industry recognizes these advantages to be cost-effective control at higher pH values, almost no loss in biocidal activity in the presence of ammonia, and effective control of bacteria, algae and mollusks.

A common way of introducing bromine based biocides into a water system is through the use of aqueous NaBr in conjunction with NaOCl bleach. The user feeds both materials to a common point whereupon the NaOCl oxidizes the bromide ion to HOBr/OBr^o. This activated solution is then introduced directly into the water system to be treated. The feeding of the two liquids in this fashion is necessary because the HOBr/OBr^o mixture is unstable and has to be generated on-site just prior to its introduction to the water. Furthermore, the feeding, and metering of two liquids is cumbersome, especially as the system has to be designed to allow time for the activation of bromide ion to occur. Consequently many biocide users have expressed the need for a single-feed, bromine-based biocide. Elemental bromine and molecular bromine chloride have been considered to meet these demands. Both are liquids at room temperature and can be fed directly to the water system, where immediate hydrolysis occurs to yield HOBr.



Properties of bromine and bromine chloride are compared in Table 1.

Table 1 - Physical Properties of Bromine and Bromine Chloride

Property	Bromine (Br ₂)	Bromine Chloride (BrCl)
Appearance	Fuming, dark-red liquid	Fuming, red liquid or gas
Boiling Point	59°C	5°C
Vapor Pressure (25°C)	214 mm	1800 mm
Corrosivity	Corrodes most metals in the presence of water	Corrodes most metals in the presence of water

It can be seen that certain characteristics of these materials--especially their corrosiveness, high vapor pressures and fuming tendencies--necessitate care and skill in their handling and use. Early efforts to overcome the deficiencies of these materials comprised complexing bromine with excess bromide ion in the presence of strong acid and stabilizing the resultant solutions with ethanolamine. The resultant solutions of ethanolammonium hydrogen perbromide contained up to 38% by weight elemental bromine. See in this connection, Favstritsky, U. S. Pat. No. 4,886,915; and Favstritsky, Hein, and Squires, U. S. Pat. No. 4,966,716.

These solutions permitted introduction of bromine to a water system using a single feed. As in the case of bromine and bromine chloride, the ethanolammonium hydrogen perbromide hydrolyzed in water to release HOBr. The vapor pressures of these solutions were lower than elemental bromine and bromine chloride. Nevertheless, the solutions still possessed measurable vapor pressures, and thus tended to produce undesirable reddish-colored vapors during storage and use.

An economically acceptable way of stabilizing high concentrations of aqueous solutions of bromine chloride is described in U.S. Pat. No. 5,141,652 to Moore, et al. The solution is prepared from bromine chloride, water and a halide salt or hydrohalic acid. These solutions were found to decompose at a rate of less than 30% per year and in cases of high halide salt concentration, less than 5% per year. Moreover, solutions containing the equivalent of 15% elemental bromine could be prepared. Unfortunately, the relatively high acidity of these solutions and their tendency to be corrosive and fuming impose limitations on their commercial acceptance.

Many solid bromine derivatives such as BCDMH (1,3-bromochloro-5,5-dimethylhydantoin) are limited in the amount of material that can be dissolved in water and fed as a liquid to the water treatment system. For example, the solubility of BCDMH in water is only around 0.15%. Another limitation of such derivatives is that at neutral pH, HOBr rapidly decomposes, eventually forming bromide ions. Thus, the ability to store and transport these aqueous solutions is greatly limited and of questionable commercial feasibility.

U.S. Pat. No. 3,558,503 to Goodenough et al. describes certain aqueous bromine solutions stabilized with various stabilizing agents and various uses to which such solutions can be put. The compositions described in the patent comprise an aqueous bromine solution having from 0.01 to 100,000 parts per million by weight of bromine values wherein the molar ratio of bromine to nitrogen present in the bromine stabilizer ranges from 2.0 to 1 to 0.5 to 1. The stabilizer used is biuret, succinimide, urea, a lower aliphatic mono- or disubstituted urea containing from 2 to 4 carbon

atoms in each substituent group, sulfamic acid, or an alkyl sulfonamide of the formula RSO_2NH_2 where R is a methyl or ethyl group. The solution also contains sufficient hydroxide additive to provide a pH in the solution ranging from 8 to 10, the hydroxide additive being an alkaline earth hydroxide or an alkali metal hydroxide.

5 U.S. Pat. No. 5,683,654 to Dallmier et al. discusses the preparation of aqueous alkali metal or alkaline earth metal hypobromite solutions by mixing an aqueous solution of alkali or alkaline earth metal hypochlorite with a water soluble bromide ion source to form a solution of unstabilized alkali or alkaline earth metal hypochlorite. To this solution is added an aqueous solution of an alkali metal sulfamate having a temperature of at least 50°C and in an amount that provides a molar ratio
10 of alkali metal sulfamate to alkali or alkaline earth metal hypobromite of from 0.5 to 6 whereby a stabilized aqueous alkali or alkaline earth metal hypobromite solution is formed. The Dallmier et al. patent teaches that much higher levels of available halogen for disinfection were attained by this approach as compared to the Goodenough et al. approach. But the Dallmier et al. patent acknowledges that in their process, the stabilization must occur quickly after the unstable NaOBr is
15 formed.

THE INVENTION

This invention involves a new process of forming concentrated aqueous solutions of biocidally active bromine and in so doing, provides novel and eminently useful concentrated aqueous biocidal solutions of bromine and bromine chloride.

20 In one of its embodiments this invention provides a process of producing a concentrated liquid biocide composition which comprises mixing (a) bromine chloride or bromine with (b) an aqueous solution of alkali metal salt of sulfamic acid (preferably the sodium salt), the solution having a pH of at least about 7, *e.g.*, in the range of 7 to about 13.5, and preferably in the range of 7 to about 12. The amounts of (a) and (b) used are such that (i) the content of active bromine in the solution
25 is at least 100,000 ppm (wt/wt) and (ii) the atom ratio of nitrogen to active bromine from (a) and (b) is greater than 1 when bromine is used, and greater than 0.93 when bromine chloride is used. It is preferred, however, to utilize an atom ratio of nitrogen to active bromine from (a) and (b) that is greater than 1 even when using bromine chloride in the process. In a preferred embodiment the aqueous solution of alkali metal salt of sulfamic acid used in the process is preformed by mixing
30 together in water, (i) sulfamic acid and/or an alkali metal salt of sulfamic acid, and (ii) alkali metal base in proportions such that an aqueous solution of alkali metal salt of sulfamic acid is formed

having a pH of at least 7. If sulfamic acid itself is used as the starting material, it is used initially as a slurry in water with which the alkali metal base is mixed.

When introducing the bromine chloride or bromine into the aqueous solution of alkali metal salt of sulfamic acid, it is desirable to maintain the desired pH of the resulting solution at 7 or above
5 by also introducing into the solution (continuously or intermittently, as desired) additional alkali metal base, such as by a co-feed of an aqueous solution of alkali metal base.

It is preferred to employ bromine chloride as the source of the active bromine in the above process because in the resulting aqueous compositions, all of the bromine of the bromine chloride is made available as active bromine in solution. In other words, the chlorine of the bromine chloride
10 is converted in the process to dissolved alkali metal chloride salt, thereby liberating the bromine as the active bromine content of the biocidal composition. Thus the more expensive component of the bromine chloride--viz., bromine -- is fully utilized in forming active bromine in the aqueous biocidal composition, and concurrently the less expensive component--the anionic chlorine in the bromine chloride -- makes this beneficial result possible.

15 By utilizing bromine or bromine chloride with caustic in the stabilized bromine composition, higher levels of active halogen are achievable, compared to the levels obtained by the addition of sodium hypochlorite to sodium bromide. The process and the compositions formed also have about twice the content of active bromine as the most concentrated solutions produced pursuant to the Goodenough, *et al.* patent. Moreover, even at the high levels of active bromine that exist in the
20 compositions of this invention, it has been found possible to provide biocidal compositions that maintain these high levels of active bromine for at least a two-month period, and that do not exhibit a visible or offensive vapor or odor during this period.

In another embodiment, alkali metal dichlorohypobromite, $M[BrCl_2]$ (M = alkali metal) is preformed by pre-mixing bromine chloride with aqueous sodium chloride, and the bromine chloride
25 is used in this form to provide the active bromine content of the resultant solution. The preferred alkali metal dichlorohypobromite is sodium dichlorohypobromite.

Another embodiment of this invention is an aqueous biocide composition comprising water having in solution therein (i) an active bromine content derived from bromine chloride of at least about 100,000 ppm (wt/wt), (ii) an alkali metal salt of sulfamic acid (preferably the sodium salt), and
30 (iii) an alkali metal chloride (preferably sodium chloride), wherein the relative proportions of (i) and (ii) are such that the atom ratio of nitrogen to active bromine is greater than 1, and wherein the pH of the composition is at least 7, e.g., in the range of 7 to about 13.5, and preferably in the range of

7 to about 12. In a less preferred embodiment (i) is bromine (Br_2) and (iii) is an alkali metal bromide (especially sodium bromide).

In each of the embodiments of this invention, the atom ratio of nitrogen to active bromine is preferably in the range of 1.1 to 1.5. Still higher ratios can be employed, if desired.

5 The above and other embodiments of this invention will be still further apparent from the ensuing description and appended claims.

A general procedure for preparing the compositions of this invention using sulfamic acid involves, as a first step, forming a slurry of sulfamic acid in water. Typically the pH of this slurry is below 1 pH unit. Sodium hydroxide at 50% concentration is then added until the solid is
10 completely dissolved. Additional 50% NaOH is added until the desired pH is reached. Bromine or bromine chloride is then added at a rate to allow the bromine to dissolve and react with the sulfamic acid without forming a pool of halogen on the bottom of the reactor. On a laboratory scale, a convenient rate of addition is approximately two drops per second. Sodium hydroxide (*e.g.*, 25% or 50%) is co-fed to the reactor to maintain the desired pH (*e.g.*, in the range of 7 to about 13.5, and
15 it may be possible to operate even at a pH in the range of 13.5 to 14. It has been found that stable solutions containing as much as 26% active bromine (11.5% on an active chlorine basis) can be prepared by the process of this invention.

The following examples are presented for purposes of illustration and not limitation.

EXAMPLES

20 Various compositions were prepared using the above general procedure and the active bromine content of the resultant compositions was determined analytically. The conditions used and results obtained (observations on odor and vapor, and initial contents of active bromine in the solutions) are summarized in Table 2.

Table 2 - Data on Prepared Sulfamic Acid Stabilized Bromine Solutions

Ex. No.	Halogen	pH	SA _{eq}	Odor and Vapor Comments	Active Br ₂ , wt%
1	Br ₂	13.0	1.42	Slight sweet smell, no observed vapor	12.4%*
2	Br ₂	7.0	1.48	Slight Br odor, no fuming	13.4%*
3**	BrCl	7	0.92	Strong Br odor, slight fuming	11.2%
4	Br ₂	13.0	1.15	Slight sweet smell, no observed vapor	19.6%
5	Br ₂	7.0	1.13	Moderate Br odor, no fuming	26.7%
6	BrCl	12.5	0.94	Slight sweet smell, no observed vapor	18.0%
7	BrCl	12.8	1.41	Slight sweet smell, no observed vapor	17.6%

SA_{eq} = Sulfamic acid to halogen mole ratio.

* Measured with Hach spectrometer; all others titrated using starch-iodine-sodium arsenite method.

** Comparative example.

The specific details for Examples 3-7 of the Table are given below. Example 8 illustrates the embodiment of the invention wherein an alkali metal dichlorohypobromite is utilized as the source of active bromine.

EXAMPLE 3

Bromine Chloride, Caustic and Sodium Sulfamate at Neutral pH

A 1 liter flask was charged with 52.0 g of sulfamic acid and 250 g of water. Sodium sulfamate was prepared by adding 60.0 g of 50% sodium hydroxide to the stirred slurry. Bromine chloride was prepared by adding 20 g of chlorine to 47.0 g of bromine. This bromine chloride was then co-fed with 210 g of 25% sodium hydroxide to maintain the pH between 6 and 8. 5 mL of 1 M Hydrochloric Acid were added to bring the final pH to approximately 7±0.5. The solution, which contained some solids, was transferred to an amber bottle for storage. Starch-iodine titration of a sample of the solution indicated that it had an active bromine concentration of 11.2%.

EXAMPLE 4

Bromine, Caustic (50% Sodium Hydroxide) and Sodium Sulfamate

A 500 mL flask was charged with 26.0g of sulfamic acid and 50 g water. To this slurry was added 35.0 g of 50% sodium hydroxide. As the acid was converted to the sodium salt, it dissolved into the aqueous solution more readily. Bromine (37.0 g) and 50% sodium hydroxide (30.0 g) were co-fed into the solution at a rate which maintained the pH between 11 and 13. After all of the

bromine and caustic had been added, the contents were transferred to an amber bottle for storage. Starch-iodine titration of a sample of the solution indicated that it had an active bromine concentration of 19.6%. Analysis of the bromine solution after 6 weeks of storage at ambient temperature indicated that it still contained more than 95% of its active bromine content.

5

EXAMPLE 5*Bromine, Caustic and Sodium Sulfamate at Neutral pH*

A 500 mL flask was charged with 26.0 g of sulfamic acid and 50 g of water. To this stirred slurry was added 30.9 g of 50% sodium hydroxide, which raised the initial pH to approximately 12. The sulfamic acid then dissolved into solution. Bromine (37.7 g) was fed into the solution until the
10 pH dropped to approximately 7, when 50% sodium hydroxide (10.9 g) was co-fed to maintain the pH between 6 and 9. 5 mL of 0.01 N sodium hydroxide was used to bring the final pH to approximately 7 ± 0.5 . The contents were then transferred to an amber bottle for storage. Starch-iodine titration of a sample of this solution indicated that it had an active bromine content of 26.7%. Analysis of the solution after six weeks of storage at ambient temperature indicated that the stabilized
15 bromine solution still contained more than 95% of its active bromine content.

EXAMPLE 6*Bromine Chloride, Caustic and Sodium Sulfamate*

A 1 liter flask was charged with 107 g of sulfamic acid and 200 g of water. Sodium sulfamate was prepared by adding 93.9 g of 50% sodium hydroxide to the stirred slurry. Bromine
20 chloride was prepared by adding 39 g of chlorine to 96.0 g of bromine. This bromine chloride was the co-fed with 319 g of 50% sodium hydroxide to maintain the pH between 11 and 13. After stirring for an additional 30 minutes, the solution, which contained some solids, was transferred to an amber bottle for storage. Starch-iodine titration of a sample of the solution indicated that it had an active bromine concentration of 18.0%. Analysis of the solution after three weeks at ambient
25 temperature indicated that the stabilized bromine solution still contained more than 90% of its active bromine content.

EXAMPLE 7*Bromine Chloride, Caustic and Sodium Sulfamate; larger scale*

A 5 liter flask was charged with 470 g of sulfamic acid and 900 g of water. Sodium sulfamate was prepared by adding 436 g of 50% sodium hydroxide to the stirred slurry. Bromine chloride was prepared by adding 120 g of chlorine to 276 g of bromine. This bromine chloride was the co-fed with 1723 g of 50% sodium hydroxide to maintain the pH between 12 and 13. After stirring for an additional 60 minutes, the orange, clear solution was transferred to a polyethylene bottle for storage. Starch-iodine titration of a sample of the solution indicated that it had an active bromine concentration of 17.6%.

EXAMPLE 8*Reducing Vapor Pressure of Sodium Dichlorohypobromite with Sodium Sulfamate*

Sodium sulfamate was prepared by slurring 24.3 g of sulfamic acid in 9 g of water. 24.0 g of 50% sodium hydroxide was added dropwise. The flask heated noticeably and the solid dissolved. This solution was dropped into 184.6 g of sodium dichlorohypobromite. Sodium dichlorohypobromite, $\text{Na}[\text{BrCl}_2]$ is prepared by adding 30.6 g of bromine chloride to 154 g of 3M aqueous sodium chloride. An additional 24 g of 50% sodium hydroxide was added to raise the pH to 7. Analysis of this solution indicated that it had an active bromine concentration of 12.0%.

Even though the claims hereinafter may refer to substances, components and/or ingredients in the present tense ("comprises," or "is"), the reference is to the substance, component or ingredient as it existed at the time just before it was first contacted, blended or mixed with one or more other substances, components and/or ingredients, or if formed in solution, as it would exist if not formed in solution, all in accordance with the present disclosure. It matters not that a substance, component or ingredient may have lost its original identity through a chemical reaction or transformation during the course of such contacting, blending, mixing, or *in situ* formation, if conducted in accordance with this disclosure.

CLAIMS

1. A process of producing a concentrated liquid biocide formulation which comprises mixing together (a) bromine chloride or bromine and (b) an aqueous solution of alkali metal salt of sulfamic acid having a pH-of at least about 7, in amounts such that (i) the active bromine content of the solution is at least about 100,000 ppm (wt/wt), and (ii) the atom ratio of nitrogen to active bromine from (a) and (b) is greater than 1 when bromine is used, and is greater than 0.93 when bromine chloride is used.
2. A process according to Claim 1 wherein (a) is bromine chloride, and wherein said atom ratio is greater than 1.
- 10 3. A process according to Claim 1 wherein said aqueous solution of alkali metal salt of sulfamic acid is an aqueous solution of the sodium salt of sulfamic acid.
4. A process according to Claim 3 wherein (a) is bromine chloride, and wherein said atom ratio is greater than 1.
5. A process according to Claim 1 wherein said aqueous solution of alkali metal salt of sulfamic acid is formed by mixing together in water, (i) sulfamic acid and/or an alkali metal salt of sulfamic acid, and (ii) alkali metal base in proportions such that an aqueous solution of alkali metal salt of sulfamic acid is formed having a pH of at least 7.
- 15 6. A process according to Claim 5 wherein (a) is bromine chloride, and wherein said atom ratio is greater than 1.
- 20 7. A process according to Claim 5 wherein said alkali metal base is a sodium base such that said aqueous solution of alkali metal salt of sulfamic acid is an aqueous solution of sodium salt of sulfamic acid.
8. A process according to Claim 7 wherein (a) is bromine chloride, and wherein said atom ratio is greater than 1.
- 25 9. A process according to Claim 5 wherein, at the time (i) and (ii) are mixed together, the alkali metal base is in the form of a preformed aqueous solution of alkali metal base, and (ii) is sulfamic acid in the form of a preformed slurry sulfamic acid in water.
10. A process according to Claim 9 wherein (a) is bromine chloride, and wherein said atom ratio is greater than 1.
- 30 11. A process according to Claim 9 wherein said preformed aqueous solution of alkali metal base is a preformed aqueous solution of sodium base.

12. A process according to Claim 11 wherein (a) is bromine chloride, and wherein said atom ratio is greater than 1.

13. A process of producing a concentrated liquid biocide formulation which comprises:

- 5 A) mixing together water, sulfamic acid and sodium hydroxide in proportions such that an aqueous solution of sodium salt of sulfamic acid is formed having a pH of at least 7;
- B) introducing bromine chloride into said aqueous solution in an amount such that (i) the active bromine content of the solution is at least about 100,000 ppm (wt/wt), and (ii) the atom ratio of nitrogen to active bromine is greater than 0.93.

10 14. A process according to Claim 13 wherein the sodium hydroxide is in the form of a preformed aqueous sodium hydroxide solution when mixed in A).

15 15. A process according to Claim 14 wherein in A) the sulfamic acid is mixed with water to form an aqueous slurry of sulfamic acid, and said preformed aqueous sodium hydroxide solution is added to said slurry.

16. A process according to Claim 13 wherein said pH is in the range of from 7 to about 13.5.

17. A process according to Claim 13 wherein said pH is in the range of from 7 to about 11.

18. A process according to Claim 13 wherein said atom ratio is greater than 1.

19. A process according to Claim 17 wherein said atom ratio is greater than 1.

20 20. A process of producing a concentrated liquid biocide formulation which comprises mixing together (a) alkali metal dichlorohypobromite and (b) an aqueous solution of alkali metal salt of sulfamic acid having a pH of at least about 7, in amounts such that (i) the active bromine content of the solution is at least about 100,000 ppm (wt/wt), and (ii) the atom ratio of nitrogen to active bromine from (a) and (b) is greater than 1.

25 21. A process according to Claim 18 wherein said alkali metal dichlorohypobromite is a preformed aqueous solution of alkali metal dichlorohypobromite, and said aqueous solution of alkali metal salt of sulfamic acid is a preformed aqueous solution of the sodium salt of sulfamic acid.

22. A process according to Claim 21 wherein said preformed aqueous solution is a preformed aqueous solution of sodium dichlorohypobromite.

30 23. A process according to Claim 22 wherein said preformed aqueous solution of sodium dichlorohypobromite is introduced into the preformed aqueous solution of the sodium salt of sulfamic acid.

24. A process according to Claim 22 wherein said pH is in the range of from 7 to about 13.5.
25. A process according to Claim 22 wherein said pH is in the range of from 7 to about 12.
- 5 26. An aqueous biocide composition comprising water having in solution therein (i) an active bromine content derived from bromine chloride of at least about 100,000 ppm (wt/wt), (ii) an alkali metal salt of sulfamic acid, and (iii) an alkali metal chloride, wherein the relative proportions of (i) and (ii) are such that the atom ratio of nitrogen to active bromine is greater than 0.93, and wherein the pH of the composition is at least 7.
- 10 27. A composition according to Claim 26 wherein said pH is in the range of from 7 to about 13.5.
28. A composition according to Claim 26 wherein said pH is in the range of from 7 to about 12.
29. A composition according to Claim 26 wherein the alkali metal salt of sulfamic acid
15 is the sodium salt of sulfamic acid, and wherein the alkali metal chloride is sodium chloride.
30. A composition according to Claim 26 wherein said atom ratio is greater than 1.
31. A composition according to Claim 28 wherein said atom ratio is greater than 1.
32. An aqueous biocide composition comprising water having in solution therein (i) an active bromine content derived from bromine of at least about 100,000 ppm (wt/wt), (ii) an alkali
20 metal salt of sulfamic acid, and (iii) an alkali metal bromide, wherein the relative proportions of (i) and (ii) are such that the atom ratio of nitrogen to active bromine is greater than 1, and wherein the pH of the composition is at least 7.
33. A composition according to Claim 32 wherein said pH is in the range of from 7 to about 13.5.
- 25 34. A composition according to Claim 32 wherein said pH is in the range of from 7 to about 12.
35. A composition according to Claim 32 wherein the alkali metal salt of sulfamic acid is the sodium salt of sulfamic acid, and wherein the alkali metal bromide is sodium bromide.

INTERNATIONAL SEARCH REPORT

International Application No

PC, /US 99/09801

A. CLASSIFICATION OF SUBJECT MATTER
IPC 6 A01N59/00 A01N25/22 C02F1/76

According to International Patent Classification (IPC) or to both national classification and IPC

B. FIELDS SEARCHED

Minimum documentation searched (classification system followed by classification symbols)

IPC 6 A01N C02F

Documentation searched other than minimum documentation to the extent that such documents are included in the fields searched

Electronic data base consulted during the international search (name of data base and, where practical, search terms used)

C. DOCUMENTS CONSIDERED TO BE RELEVANT

Category *	Citation of document, with indication, where appropriate, of the relevant passages	Relevant to claim No.
A	US 3 558 503 A (GOODENOUGH ROBERT D ET AL) 26 January 1971 (1971-01-26) cited in the application	1-35
X	column 1, line 64 - column 2, line 22 column 2, line 28 - line 51 column 2, line 55 - line 64 column 4, line 66 - line 76; claims	26-35
A	WO 97 34827 A (NALCO CHEMICAL CO) 25 September 1997 (1997-09-25) page 1 - page 4, paragraph 1	1-35
X	page 5, last paragraph - page 7, line 11; claims; example 1 & US 5 683 654 A (DALLMIER) cited in the application	26-31
	-/--	

☒ Further documents are listed in the continuation of box C.

☒ Patent family members are listed in annex.

* Special categories of cited documents:

"A" document defining the general state of the art which is not considered to be of particular relevance

"E" earlier document but published on or after the international filing date

"L" document which may throw doubts on priority claim(s) or which is cited to establish the publication date of another citation or other special reason (as specified)

"O" document referring to an oral disclosure, use, exhibition or other means

"P" document published prior to the international filing date but later than the priority date claimed

"T" later document published after the international filing date or priority date and not in conflict with the application but cited to understand the principle or theory underlying the invention

"X" document of particular relevance; the claimed invention cannot be considered novel or cannot be considered to involve an inventive step when the document is taken alone

"Y" document of particular relevance; the claimed invention cannot be considered to involve an inventive step when the document is combined with one or more other such documents, such combination being obvious to a person skilled in the art.

"&" document member of the same patent family

Date of the actual completion of the international search

2 September 1999

Date of mailing of the international search report

15/09/1999

Name and mailing address of the ISA

European Patent Office, P.B. 5818 Patentlaan 2
NL - 2280 HV Rijswijk
Tel. (+31-70) 340-2040, Tx. 31 651 epo nl.
Fax: (+31-70) 340-3016

Authorized officer

Muellners, W

INTERNATIONAL SEARCH REPORT

International Application No

PC., US 99/09801

C.(Continuation) DOCUMENTS CONSIDERED TO BE RELEVANT

Category	Citation of document, with indication, where appropriate, of the relevant passages	Relevant to claim No.
A	US 5 141 652 A (MOORE JR ROBERT M ET AL) 25 August 1992 (1992-08-25) cited in the application column 1 - column 2, line 31 ---	1-35
A	US 3 152 073 A (J.R.MORTON) 6 October 1964 (1964-10-06) column 3, line 52 - column 4, line 55 column 4, line 74 - column 5, line 21; claims 4,5 table II -----	20-25

INTERNATIONAL SEARCH REPORT

Information on patent family members

International Application No

PCT, JS 99/09801

Patent document cited in search report		Publication date	Patent family member(s)	Publication date
US 3558503	A	26-01-1971	NONE	
WO 9734827	A	25-09-1997	US 5683654 A	04-11-1997
			US 5795487 A	18-08-1998
			AU 2603397 A	10-10-1997
			BR 9702226 A	23-02-1999
			CA 2221927 A	25-09-1997
			EP 0827486 A	11-03-1998
			NO 971346 A	23-09-1997
US 5141652	A	25-08-1992	DE 69212244 D	22-08-1996
			DE 69212244 T	21-11-1996
			EP 0518119 A	16-12-1992
US 3152073	A	06-10-1964	NONE	